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17-4 17.16 The  $Q$  for the original reaction is  $Q_{ref} = 2.22 \times 10^2$   $\frac{HS}{HS}$  a) The given reaction  $\frac{1}{2} S_2(g) + H_2(g) \rightleftharpoons H_2S(g)$  is the reverse reaction of the original reaction multiplied by a factor of  $\frac{1}{2}$ . The equilibrium constant for the reverse reaction is the inverse of the original constant.

**CHAPTER 17 EQUILIBRIUM: THE EXTENT OF CHEMICAL REACTIONS**

Ch 17 Thermochemistry Practice Test Matching Match each item with the correct statement below. a. calorimeter d. enthalpy b. calorie e. specific heat c. joule f. heat capacity \_\_\_\_\_ 1. quantity of heat needed to raise the temperature of 1 g of water by  $1^\circ C$  \_\_\_\_\_ 2. SI unit of energy

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